Chemistry Content Standards:
   The kinetic molecular theory describes the motion of atoms and molecules and explains the properties of gases.
14.3.1   Compute the value of an unknown using the ideal gas law.
14.3.2   Compare and contrast real and ideal gases.

Stamp: Trans # 10
                 Page 419, 7-8 (Boyles)
                 Page 421, 9-10 (Charles)
                 Page 417, 1-6

Lecture: Gases 14.3 PPT Ideal Gas Law page 54 in NB

Chemistry

               Ideal: PV=nRT
                 n=number of moles
                 R= constant (atm or kPa)
                     R: atm=0.0821 L x atm / mol x K
                     R: kPa=8.315   L x kPa / mol x K

Classwork: WB 14.3
                 Ideal Gas Laws Page 152-153
                 14.2 Section Review, B-D
                 Gas Laws Practice Problems

Homework: Page 429, 27-30